

**POLYATOMIC IONS**

	<b>-1 anions</b>		<b>-2 anions</b>
hydroxide	$\text{OH}^-$	sulfate	$\text{SO}_4^{2-}$
nitrate	$\text{NO}_3^-$	sulfite	$\text{SO}_3^{2-}$
nitrite	$\text{NO}_2^-$	carbonate	$\text{CO}_3^{2-}$
cyanide	$\text{CN}^-$	chromate	$\text{CrO}_4^{2-}$
thiocyanate	$\text{SCN}^-$	dichromate	$\text{Cr}_2\text{O}_7^{2-}$
bicarbonate	$\text{HCO}_3^-$	hydrogen phosphate	$\text{HPO}_4^{2-}$
bisulfate	$\text{HSO}_4^-$	peroxide	$\text{O}_2^{2-}$
acetate (organic)	$\text{CH}_3\text{COO}^-$		
acetate (alternate)	$\text{C}_2\text{H}_3\text{O}_2^-$		<b>-3 anions</b>
dihydrogen phosphate	$\text{H}_2\text{PO}_4^-$	phosphate	$\text{PO}_4^{3-}$
permanganate	$\text{MnO}_4^-$		
hypochlorite	$\text{ClO}^-$		
chlorite	$\text{ClO}_2^-$		<b>cations</b>
chlorate	$\text{ClO}_3^-$	Ammonium	$\text{NH}_4^+$
perchlorate	$\text{ClO}_4^-$	Mercury(I)	$\text{Hg}_2^{2+}$

# NOMENCLATURE FLOWCHART

## Flow Chart for Naming Simple Inorganic Compounds

The flowchart is adapted from p. 131-132 of the February 1983 issue of the *Journal of Chemical Education*.

